

Recall...

Symbol

A one- or two-letter abbreviation derived from the element's English or Latin name.

Name

Element's common name.

Mass Number

The sum of the numbers of protons and neutrons in a specific isotope.

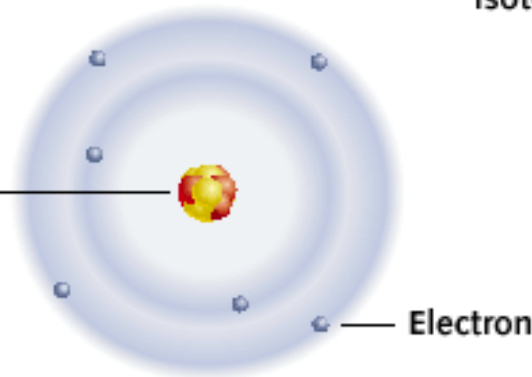
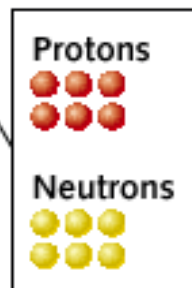
Atomic Number

Equal to the number of protons in the nucleus, as well as the number of electrons in the electron cloud.

Relative Atomic Mass

Weighted average of the masses of all the element's isotopes. Rounding the atomic mass to the nearest whole number yields the mass number of the most common isotope.

6
C
Carbon
12.011



Carbon Atom

Isotopes- each of two or more forms of the same element that contain different numbers of neutrons

Name that Isotope!

Write “name that isotope #_____” in your notebook

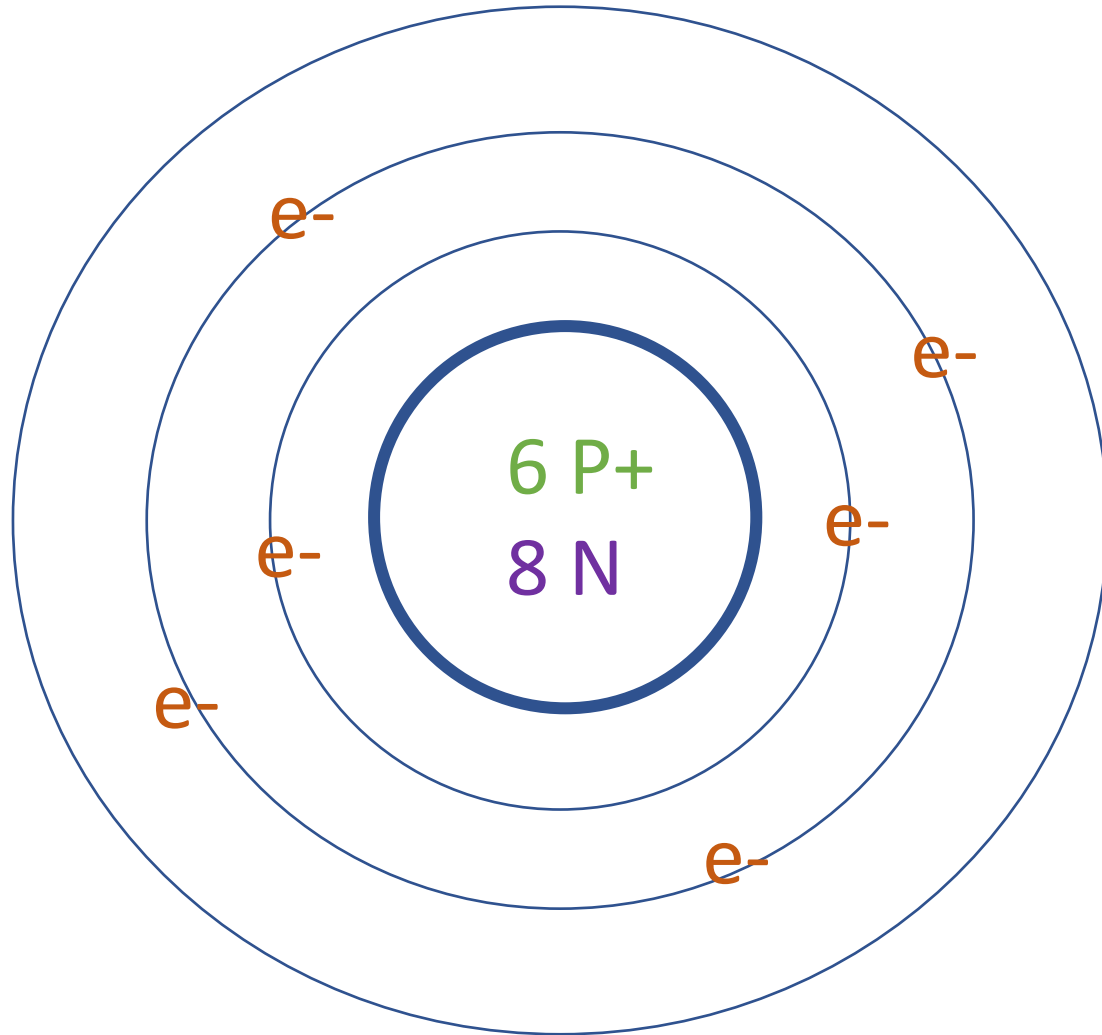
Write the element name followed by a dash and the mass number of the isotope

Example:

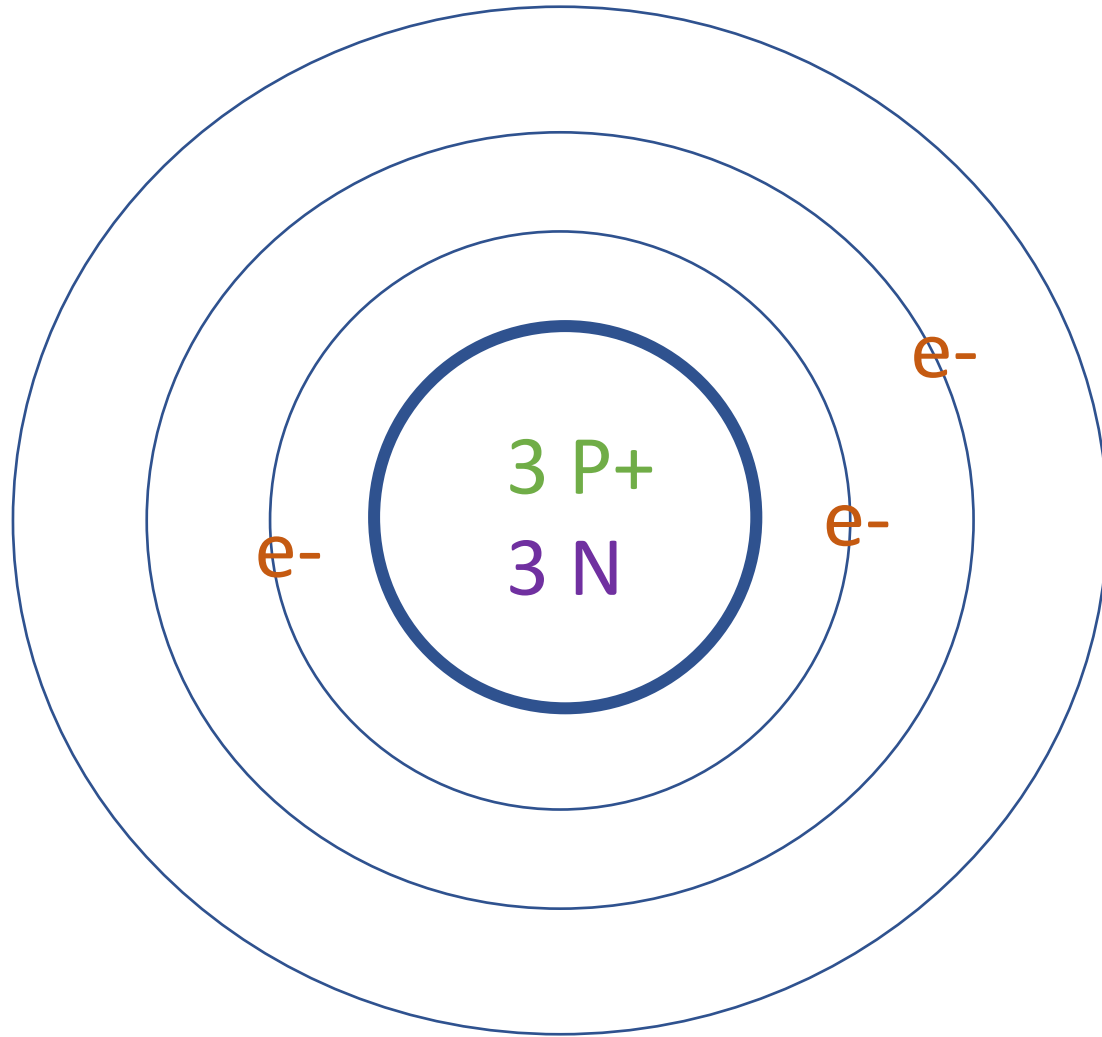
Name that Isotope # 487

Molybdenum- 95

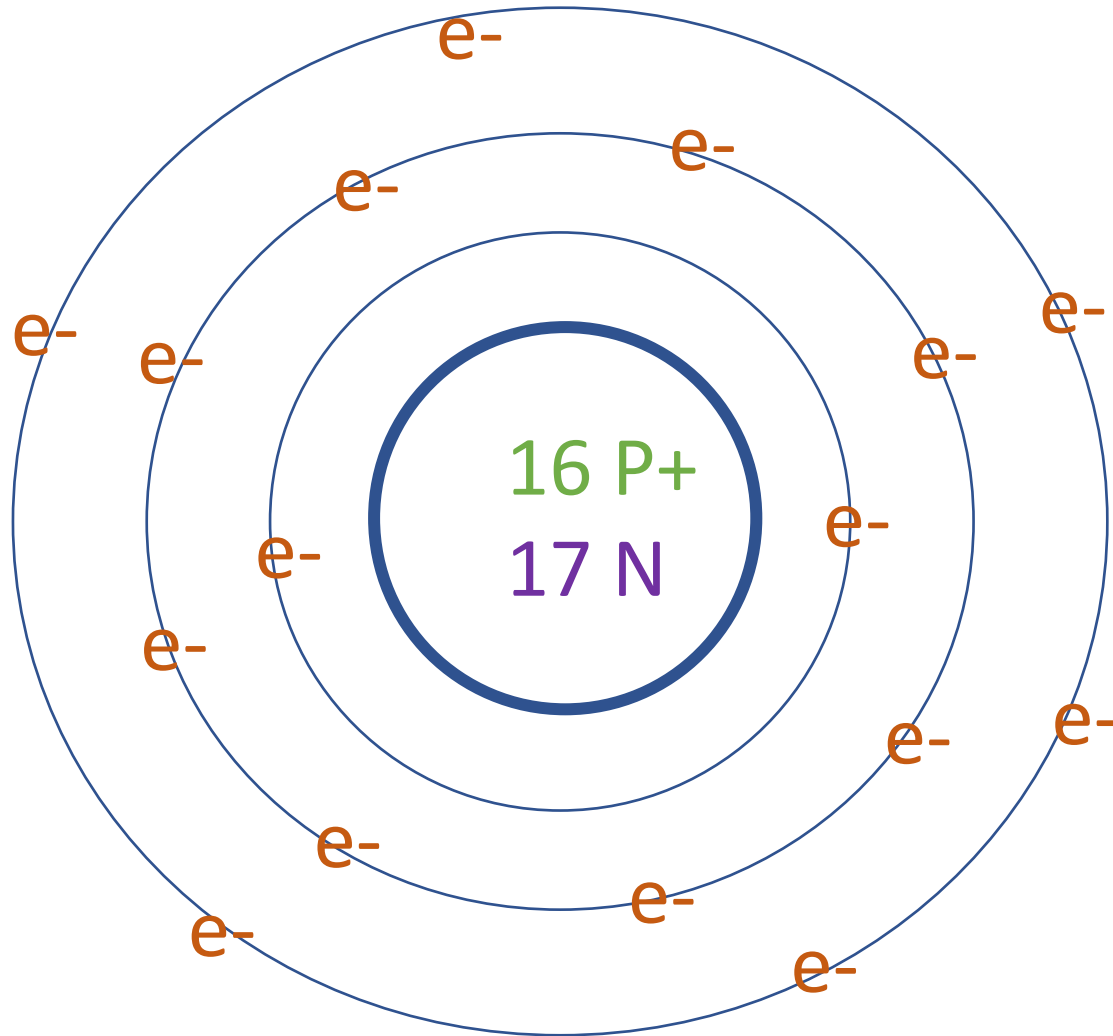
Name that Isotope #1



Name that Isotope #2



Name that Isotope #3



Name that Ion!

Write “Name that ion #_____” in your notebook

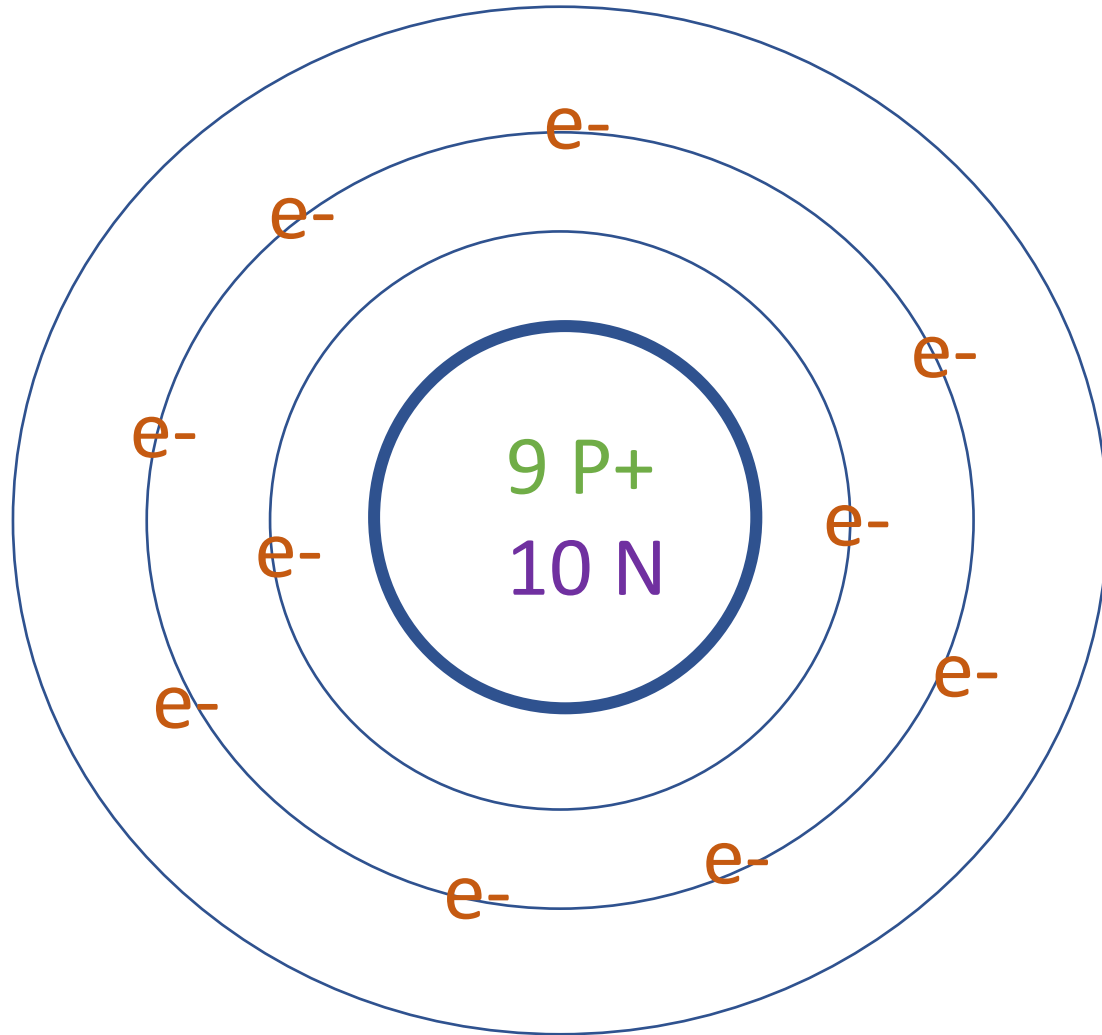
Write the element symbol followed its charge written as a superscript (small and up to the right)

Example:

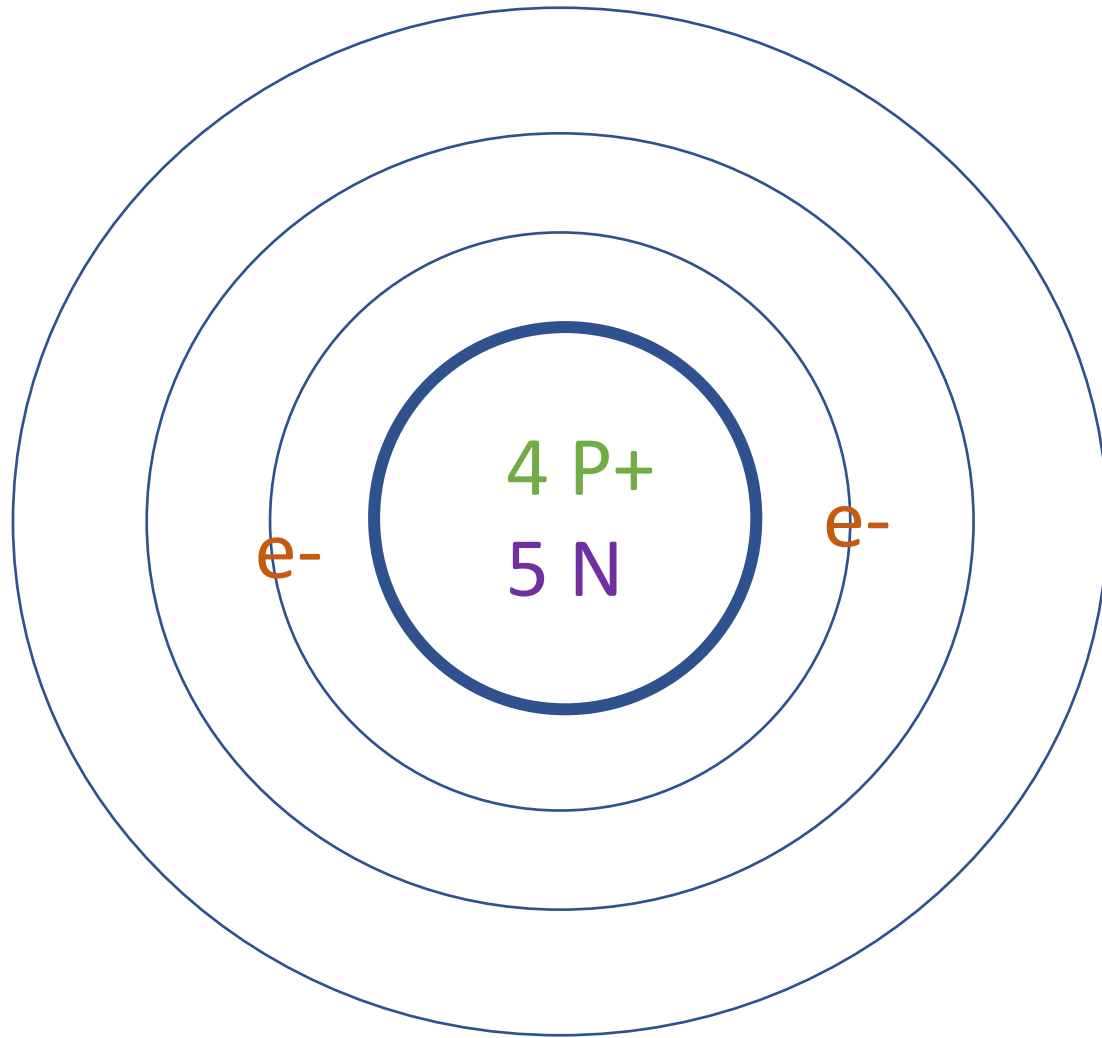
Name that Ion # 23,472



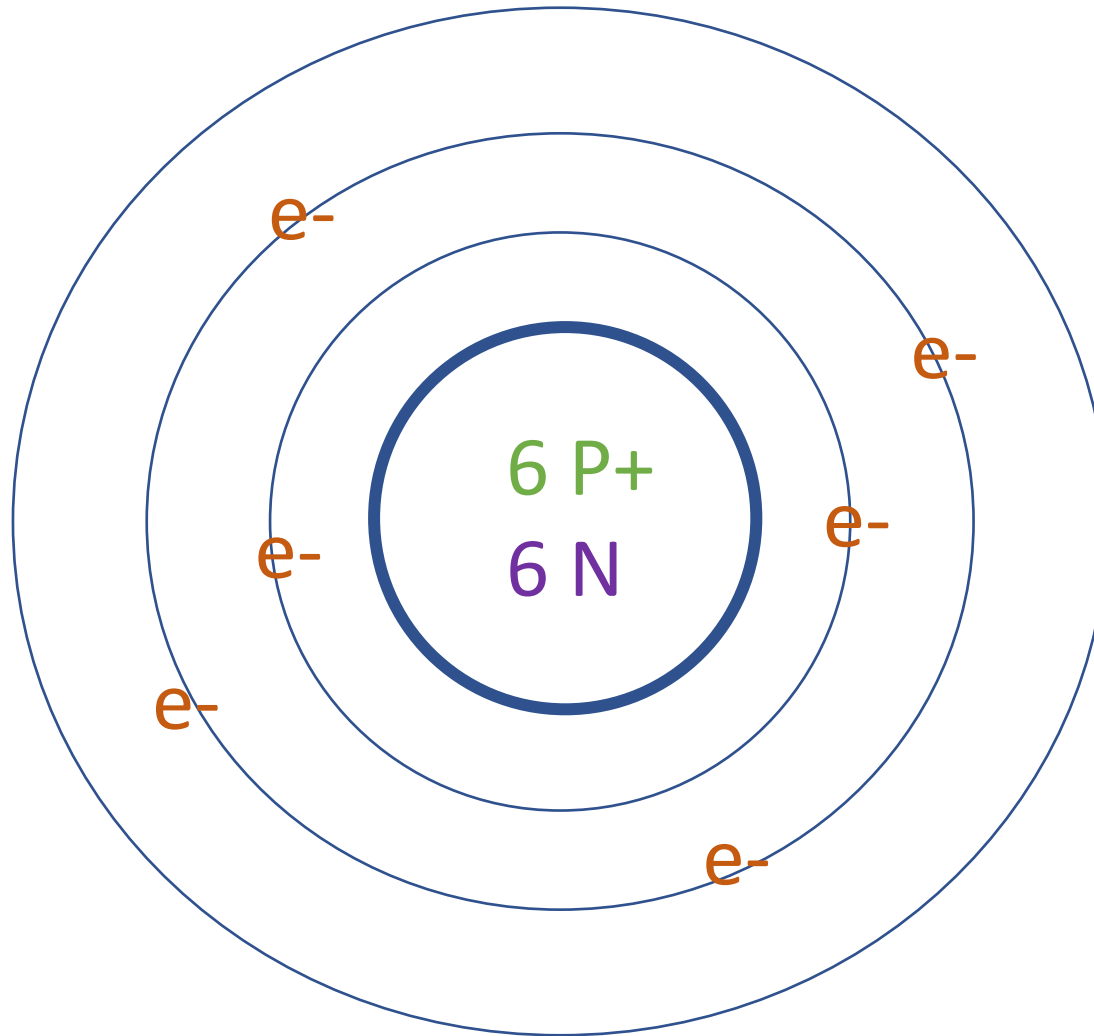
Name that Ion #1



Name that Ion #2



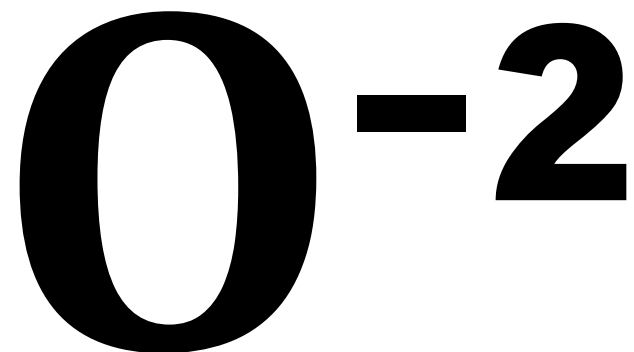
Name that Ion #3



Draw That Ion!

Draw the Ion

Draw that Ion #1



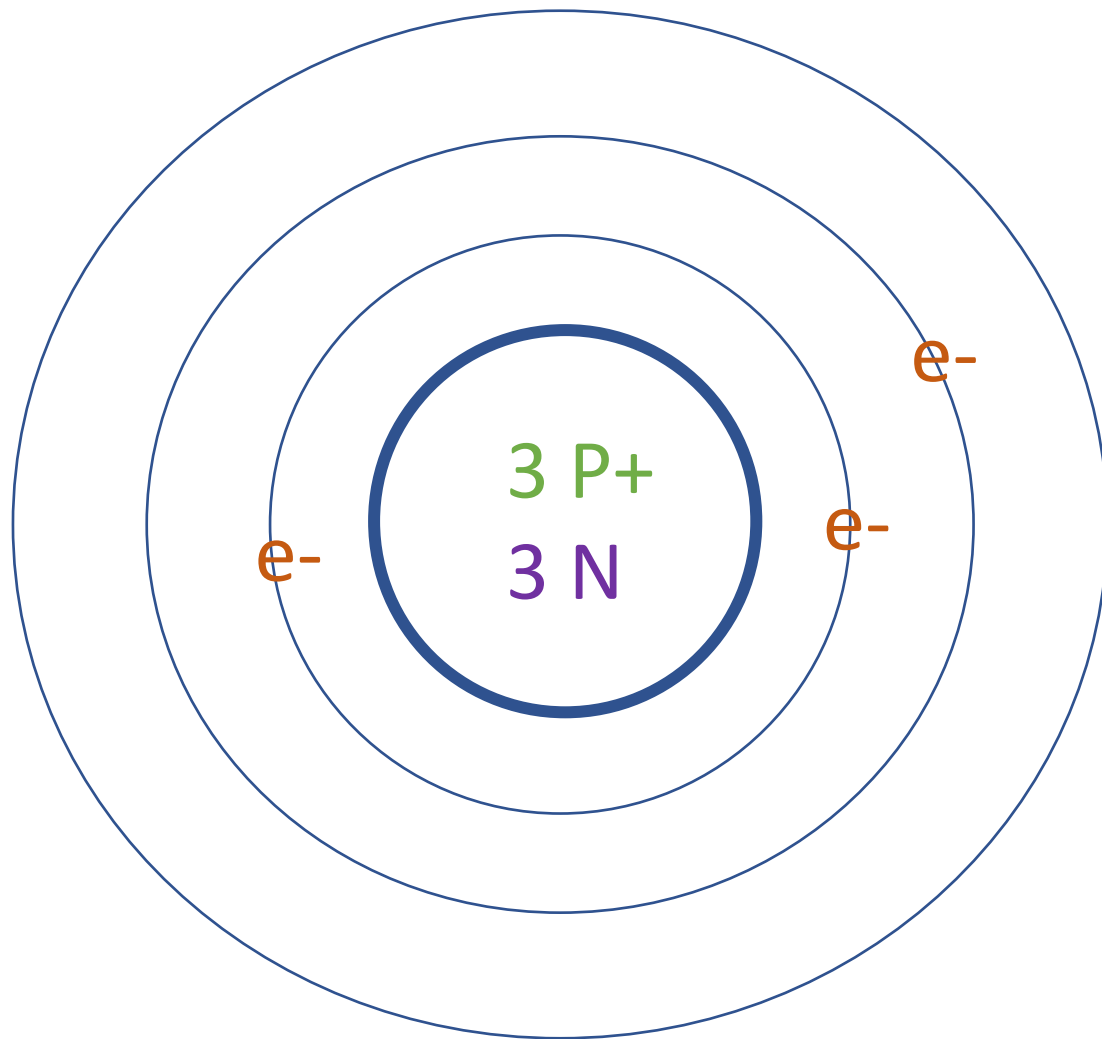
Draw that Ion #2

Cl⁻

Isotope Notation



Write the Isotope Notation



Write the Isotope Notation..

